

Pergamon

Tetrahedron Letters, Vol. 35, No. 28, pp. 5019-5022, 1994 Elsevier Science Ltd Printed in Great Britain 0040-4039/94 \$7.00+0.00

0040-4039(94)01015-3

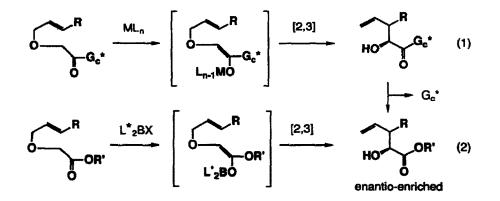
Enantioselective [2,3]Wittig Rearrangement Involving a Chiral Boron Enolate Terminus

Katsuhiko Fujimoto¹ and Takeshi Nakai*

Department of Chemical Technology, Tokyo Institute of Technology Meguro-ku, Tokyo 152, Japan

Abstract: The first enantioselective version of the ester enolate [2,3]Wittig rearrangement is described which involves a chiral boron enolate with a chiral bis-sulfonamide ligand to provide a high enantioselectivity(>95%ee), along with a high three diastereoselectivity.

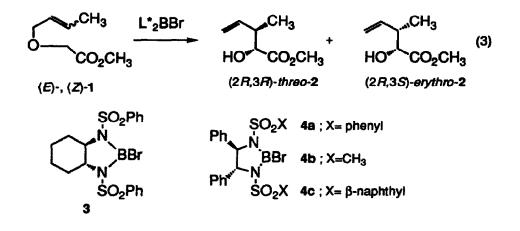
The [2,3]Wittig sigmatropic rearrangement enjoys wide application in organic synthesis.² Of special value among many variants is the "enolate [2,3]Wittig" rearrangement that involves an enolate as the migrating terminus to eventually provide α -hydroxy- β -alkyl carboxylic acid derivatives of biological and synthetic importance.² Recently several asymmetric versions of the enolate [2,3]Wittig process have been developed which involve a chiral enolate terminus generated from optically active substrates or auxiliaries (eq 1).³ However, no example has been reported of the *enantioselective* version involving a chiral ligand-bound ester enolate.⁴ Herein we wish to report the first example of the enantioselective version of the ester enolate [2,3]Wittig process (eq 2).⁵ The key to this success is the use of a chiral boron enolate terminus containing a chiral bis-sulfonamide as the controller ligand (L*).



After several attempts,⁶ we chose as a chiral ligand on boron chiral bis-sulfonamides, easily prepared from

commercially available (1R,2R)-1,2-diaminocyclohexane^{7a} and (1R,2R)-1,2-diamino-1,2-diphenylethane^{7b}, and prepared the chiral boron reagents 3 and 4 from BBr₃ and the chiral bis-sulfonamides according to the literature procedure.^{7c}

First, we studied the rearrangement of (*E*)-1 using the boron reagents 3 and 4 (eq 3). Thus, (*E*)-1 was treated with a slight excess of 3 or 4 in the presence of Et₃N in CH₂Cl₂ at -50 °C to afford a diastereomeric mixture of the enantio-enriched α -hydroxy- β -methyl ester 2 in moderate yields.⁸ The results are summarized in Table 1 (entries 1-4).



As seen in Table 1, the present rearrangements uniformly exhibit (R)-preference at C-2 of 2 and three preference between the two new chiral centers. Interestingly, the yields and the levels of enantio- and diastereoselection vary critically with the structure of chiral ligands. While the use of 3 leads to the modest levels of enantioselectivity (62%ee) and three selectivity (68%), the use of 4a provides the highest degrees of enantioselectivity (96%ee) and three selectivity (83%). Of particular interest is the relatively high $E \rightarrow three$ selection observed here, which is in sharp contrast to the high $E \rightarrow erythree$ selections previously reported for the [2,3]Wittig rearrangements of (E)-1 involving lithium,⁹ zirconium,¹⁰ tin,¹¹ titanium,¹¹ and dialkylboron ¹² enolates.

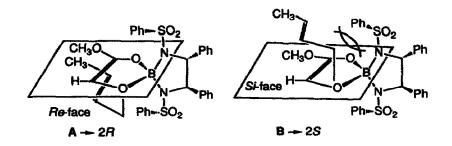
Next, we turned our attention to the rearrangement of (Z)-1. Under the same conditions as described above, (Z)-1 was treated with 4a to afford a stereoisomeric mixture of 2, again, in a high enantioselectivity (94%ee) and *threo* selectivity (84%) (entry 5). These findings indicate that, surprisingly enough, the present rearrangements of (E)- and (Z)-1 exhibit the identical senses of both enantio- and diastereoselection regardless of the crotyl geometry, suggesting that the stereo-directing power of the chiral ligand is much stronger than that of the substrate geometry.

Although we have no definitive explanation at present for the unusual *threo* selection owing to the great complexities of this process, the observed sense of enantioselection (2R) might be interpreted as follows. The enolization leads to the metal-chelated (*E*)-enolate which undergoes the [2,3]-shift preferentially from the bottom side (*re*-face) as depicted in formura A, where the steric repulsion between the crotyl and phenyl groups would be minimized.

Entry	Substrate	L*2BBr	%Yield	threo:erythro a	%ee b	Config.c
1	(<i>E</i>)-1	3	69	68:32	62 (<i>threo</i>)	2R,3R
					40 (<i>erythro</i>)	2R,3S
2	(<i>E</i>)-1	4a	66	83:17	96 (<i>threo</i>)	2R,3R
					71 (<i>erythro</i>)	2R,3S
3	(<i>E</i>)-1	4 b	57	63:37	90 (<i>threo</i>)	2R,3R
					68 (<i>erythro</i>)	2R,3S
4	(<i>E</i>)-1	4 c	27	66:34	90 (<i>threo</i>)	2R,3R
					55 (erythro)	2R,3S
5	(<i>Z</i>)-1	4 a	56	84:16	94 (<i>threo</i>)	2R,3R
					38 (erythro)	2R,3S

Table 1. Eantioselective Rearrangements of (E)- or (Z)-1

(a) Determined by ¹HNMR analysis. (b) Determined by HPLC analysis of the MTPA esters (see ref 13). (c) Assigned by LIS-NMR analysis as described in our previous paper (ref 3a).



In summary, we have developed the first enantioselective version of the ester enolate [2,3]Wittig rearrangement which involves a chiral ligand-bound boron enolate to provide the α -hydroxy- β -methyl ester in a high enantio- and *threo*-selectivity. Futher efforts are in progress to prove this interesting process and to apply the present methodology to natural product synthesis.

Acknowledgment. This work was supported in part by a Grant-in-Aid for Scientific Research from the Ministry of Education, Science and Culture, Japan.

REFERENCE AND NOTES

- 1. A visiting research fellow from Sankyo Co., Tokyo.
- Reviews: Nakai, T.; Mikami, K. Chem. Rev. 1986, 86, 885. Mikami, K.; Nakai, T. Synthesis 1991, 594. Nakai, T.; Mikami, K. Org. React. in press

- (a) Mikami, K.; Fujimoto, K.; Kasuga, T.; Nakai, T. Tetrahedron Lett. 1984, 25, 6011. (b) Mikami, K.; Kasuga, T.; Fujimoto, K.; Nakai, T. Ibid. 1986, 27, 4185. (c) Mikami, K.; Takahashi, O.; Kasuga, T.; Nakai, T. Chem. Lett. 1985, 1729. (d) Takahashi, O.; Mikami, K.; Nakai, T. Ibid. 1987, 69. (e) Uchikawa, M.; Hanamoto, T.; Katsuki, T.; Yamaguchi, M. Tetrahedron Lett. 1986, 27, 4577.
- For the enantioselective versions of different [2,3]Wittig variants (25-70%ee) with chiral lithium amide bases: Marshall, J. A.; Lebreton, J. J. Am. Chem. Soc. 1988, 110, 2925. Idem. Tetrahedron Lett. 1987, 28, 3323. Idem. J. Org. Chem. 1988, 53, 4108. Marshall, J. A.; Wang, X. J. Ibid. 1990, 55, 2995.
- Part of this work was presented at the 65th Annual Meeting of the Chemical Society of Japan, Tokyo, 1993.
- 6. Initial attempts using [(+)-Ipc]₂BX (X=Cl, Br, I, OTf) failed; no [2,3]Wittig product was obtained.
- (a) Takahashi, H.; Kawakita T.; Yoshioka, M.; Kobayashi, S.; Ohno, M. Tetrahedron Lett. 1989, 30, 7095. (b) Corey, E. J.; Imwinkelried, R.; Pikul, S.; Xiang, Y. B. J. Am. Chem. Soc. 1989, 111, 5493. (c) Corey, E. J.; Kim, S. S. J. Am. Chem. Soc. 1990, 112, 4976.
- 8. Typical Experimental Procedure: To a solution of 4a [prepared from (1R,2R)-1,2-N,N'-bis(benzenesulfonylamino)-1,2-diphenylethane (0.64g, 1.3 mmol)] in CH₂Cl₂ (3 mL) was added dropwise a solution of (E)-1 (0.144g, 1 mmol) and Et₃N (0.21 mL, 1.5 mmol) in CH₂Cl₂ (2.5 mL) at -50 °C. The resulting mixture was stirred for 2 h at -50 °C and then quenched by addition of water at -50 °C. The organic layer was separated and the aqueous layer was extracted with CH₂Cl₂. The combined extracts were dried over MgSO₄, filtered, and concentrated under reduced pressure to afford the crude product. Purification of the crude product by silica gel chromatography (hexane:ethyl acetate=13:1) gave 2 (0.095g, 66% yield).
- 9. Takahashi, O.; Saka, T.; Mikami, K.; Nakai, T. Chem. Lett. 1986,1599.
- 10. Uchikawa, M.; Katsuki, T.; Yamaguchi, M. Tetrahedron Lett. 1986, 27, 4581.
- 11. Mikami, K.; Takahashi, O.; Fujimoto, K.; Nakai, T. Synlett 1991, 629.
- (a) Oh, T.; Wrobe, Z.; Rubenstein, S. M. Tetrahedron Lett. 1991, 32, 4647 [rearrangement with (n-Bu)₂BOTf]. (b) Our unpublished result (ref 5). The rearrangement with 9-BBNOTf was found to exhibit 95% of erythro selection.
- 13. (R)-MTPA esters of 2 : HPLC [L-column ODS ; eluent , CH₃CN : H₂O = 2 : 1 (0.01M AcONH₄) ; flow rate 1.0 mL/min ; detection, 254 nm] : 2*S*,3*S* isomer, t_R=14.7 min ; 2*S*,3*R* isomer, t_R=15.2 min ; 2*R*,3*R* isomer, t_R=16.0 min ; 2*R*,3*S* isomer, t_R=16.6 min.

(Received in Japan 12 April 1994)